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 $2\text{ KClO}_3 \rightarrow 2\text{ KCl} + 3\text{ O}_2$   
 $2\text{ NaCl} + \text{F}_2 \rightarrow 2\text{ NaF} + \text{Cl}_2$   
 $4\text{ H}_2 + 2\text{ O}_2 \rightarrow 2\text{ H}_2\text{O}$   
 $5\text{ Pb} + 2\text{ OH}^- + 2\text{ HCl} \rightarrow 2\text{ H}_2\text{O} + 2\text{ PbCl}_2$   
 $6\text{ Al} + 3\text{ Br}_2 + 3\text{ K}_2\text{SO}_4 \rightarrow 4\text{ KBr} + \text{Al}_2\text{SO}_4 + 3\text{ H}_2\text{O}$   
 $1\text{ C}_3\text{H}_8 + 5\text{ O}_2 \rightarrow 3\text{ CO}_2 + 4\text{ H}_2\text{O}$

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following equations  $1 \text{ N}_2 + 3 \text{ F}_2 \rightarrow 2 \text{ NF}_3$   $2 \text{ C}_6\text{H}_{10} + 17 \text{ O}_2 \rightarrow 12 \text{ CO}_2 + 10 \text{ H}_2\text{O}$   $3 \text{ HBr} + \text{KHCO}_3 \rightarrow \text{H}_2\text{O} + \text{KBr} + \text{CO}_2$   $\text{GaBr}_3 + 3 \text{ Na}_2\text{SO}_3 \rightarrow \text{Ga}_2\text{SO}_3 + 3 \text{ NaBr} + 3 \text{ SO}_2$   $5 \text{ NF}_3 + 2 \text{ N}_2\text{O}_3 \rightarrow 3 \text{ N}_2 + 5 \text{ O}_2$  using the equation from problem 2 above answer the following

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